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Solid State Physics Chapter 1: Crystal Structure

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The focus of this chapter is on one-dimensional periodic potentials, with traditional contributions going back to the early times of the foundation of quantum mechanics; brief digressions are given to evidence the conceptual link between this historical area of solid state and other much younger and very active areas of research such as artificial structures and superlattices, photonic crystals, atomic condensates.

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If NaCl is doped with 10 – 4 mol % of Sr Cl 2, the concentration of cation vacancies will be (N A = 6.02 × 10 23 mol –1) [CBSE AIPMT 2007] A 6.02 × 1016 mol –1

Solid State MCQs | Question – Paper 1

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Metallic solid conducts electricity in solid state but ionic solids do so only in molten state or in solution or metals conduct electricity through electrons and ionic substances through ions. Metallic solids are malleable and ductile while ionic solids are hard and brittle.

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Ans: It is a covalent or network solid. 1.8. Why are ionic solids conducting in the molten state and not in the solid state ? Ans: In the ionic solids, the electrical conductivity is due to the movement of the ions. Since the ionic mobility is negligible in the solid state, these are non-conducting in this state.

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